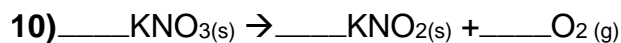
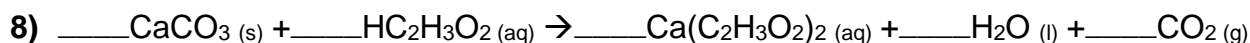
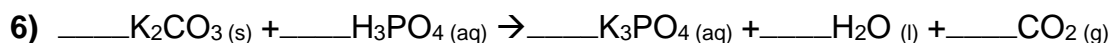
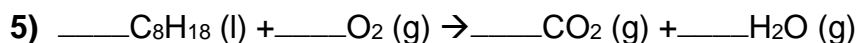
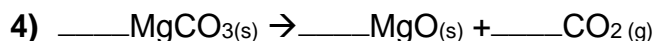
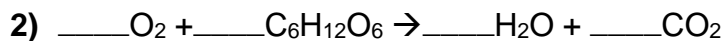
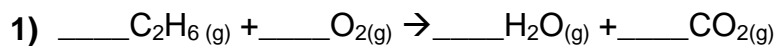


Name: _____

Period: _____

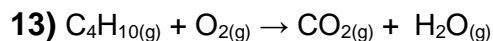
Seat#: _____



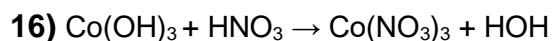
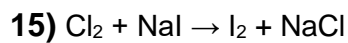
Write the balanced equation for the following and identify the type of reaction:

11) Magnesium Chloride plus Oxygen yield Magnesium Chlorate

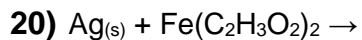
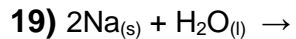
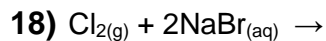
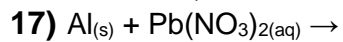
12) Silver Oxide \rightarrow Silver and Oxygen



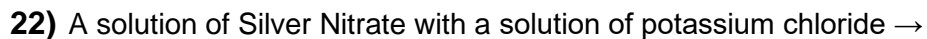
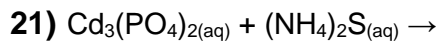
14) Sodium chloride comes apart

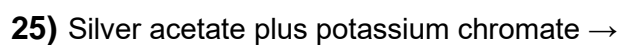
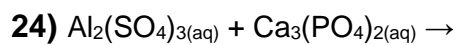
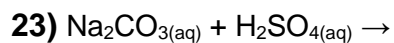


Single Replacement (Must write NET IONIC – assume all compounds are in solution (aqueous) unless otherwise stated, use Activity Series to determine if reaction takes place). Put a BOX around Net Ionic.



Double Replacement (Must write NET IONIC – assume all reactant are aqueous unless otherwise stated). Put a BOX around Net Ionic.





Determine if a precipitate forms when the following solutions are mixed. Write the overall equation, the total ionic equation, and the net ionic equation. (Double Replacement). Put a BOX around Net Ionic.

