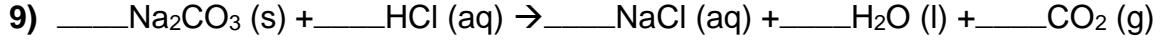
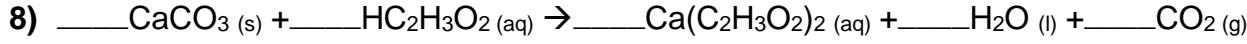
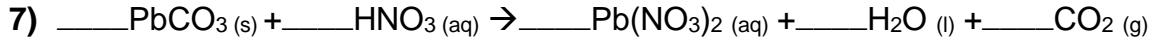
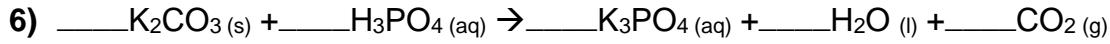
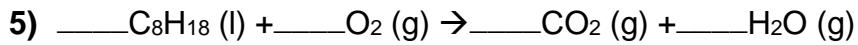
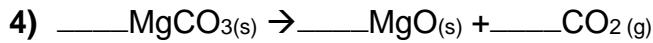
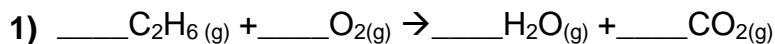


Name: \_\_\_\_\_

Period: \_\_\_\_\_

Seat#: \_\_\_\_\_



**Write the balanced equation for the following and identify the type of reaction:**

11) Magnesium Chloride plus Oxygen yield Magnesium Chlorate

12) Silver Oxide → Silver and Oxygen

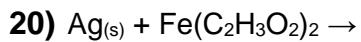
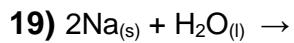
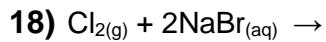
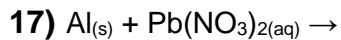
13)  $\text{C}_4\text{H}_{10\text{(g)}} + \text{O}_{2\text{(g)}} \rightarrow \text{CO}_{2\text{(g)}} + \text{H}_2\text{O}_{\text{(g)}}$

14) Sodium chloride comes apart

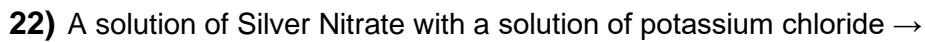
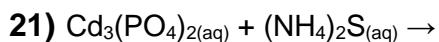
15)  $\text{Cl}_2 + \text{NaI} \rightarrow \text{I}_2 + \text{NaCl}$

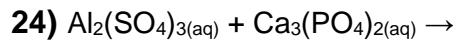
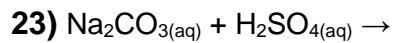
16)  $\text{Co(OH)}_3 + \text{HNO}_3 \rightarrow \text{Co(NO}_3)_3 + \text{HOH}$

**Single Replacement (Must write NET IONIC – assume all compounds are in solution (aqueous) unless otherwise stated, use Activity Series to determine if reaction takes place). Put a BOX around Net Ionic.**



**Double Replacement (Must write NET IONIC – assume all reactant are aqueous unless otherwise stated). Put a BOX around Net Ionic.**





Determine if a precipitate forms when the following solutions are mixed. Write the overall equation, the total ionic equation, and the net ionic equation. (Double Replacement). Put a BOX around Net Ionic.

**26)** A solution of ammonium carbonate is mixed with a solution of calcium acetate

**27)** A solution of sodium chromate is mixed with a solution of barium sulfate